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when a stress is applied to a system in dynamic equilibrium, the system changes in a way that relieves the stress (18.2) equilibrium constant (K_{eq}) the ratio of product concentrations to reactant concentrations at equilibrium, with each concentration raised to a power equal to the number of moles of that substance in the balanced chemical equation (18.2)

Chapter 18: Chemical Equilibrium

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562 Chapter 18 Chemical Equilibrium Figure 18-3 If the only way in or out of San Francisco and Sausalito is across the Golden Gate Bridge, which joins the two cities, then the number of vehicles in the two cities will remain constant if the number of vehicles per hour crossing the bridge in one direction equals the number of vehi-

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Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the K_{sp} for that solution, listed on the left. _____ The ion product exceeds the K_{sp} .

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Section 3 Answers

expression for each of several chemical equations that represent equilibrium situations. Homework Daily Homework: Assign 18-2 and 18-3. Section Review: p. 559. Assign items 2– 9. Chapter Review: p. 586. Assign items 2– 3 and 24– 27. Practice: p. 559. Assign problems 1– 4. ChemFile Mini-Guide to Problem Solving: Chapter 20: Equilibrium

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Chapter 18 Review Chemical Equilibrium

sometimes referred to as the chemical-equilibrium expression. The H₂, I₂, HI Equilibrium System Consider the reaction between H₂ and I₂ vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing the rate at which the violet color of the iodine vapor diminishes, as [C]^x[D]^y [A]ⁿ[B]^m 556 CHAPTER 18

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